**Unit 2 Describing Matter Vocabulary**

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| **atoms**   (atuhmz) *noun*    the basic building blocks of all matter |

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|  | **chemical property**        a characteristic that describes how a substance will interact with other substances during a chemical reaction |

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|  | **chemistry**   (kEHmuhstree) *noun*    the study of the property of matter and how matter changes |

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|  | **compound**   (kompound) *noun,verb,adjective*    a pure substance that forms when two or more elements join chemically |

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|  | **density**   (dEHnsuhtee)     physical property that compares the mass of a substance per unit per volume |

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|  | **element**   (EHluhmuhnt) *noun*    a pure substance that cannot be broken down into simpler substances by any ordinary chemical means |

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|  | **heterogeneous mixture**        pure substances are unevenly distributed throughout the mixture |

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|  | **homogenous mixture**        pure substances are evenly distributed throughout the mixture |

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|  | **mass**   (mas)     measure of the amount of matter making up an object |

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|  | **matter**   (matur) *noun,verb*    anything that has mass and takes up space |

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|  | **mixture**   (mihkschur)     type of matter that forms when two or more substances combine without joining together chemically |

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|  | **molecule**   (moluhkyool) *noun*    two or more atoms that are bonded together chemically; smallest unit of a compound that has all the properties of the compound |

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|  | **physical property**        a characteristic that can be observed or measured without changing the identity of the substance |

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|  | **solution**   (suhlooshuhn)     a well-mixed mixture containing a solvent and at least one solute that has the same properties throughout |

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|  | **substance**   (suhbstuhns) *noun*    a single kind of matter |

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|  | **volume**   (volyoom)     the amount of space matter occupies |

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|  | **weight**   (wayt)     a measure of the pull of gravity on the mass of an object |

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| **Unit 2 States of Matter Vocabulary**  **Law of Conservation of Energy**        energy can neither be created nor destroyed |

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|  | **Law of Conservation of Mass**        the total amount of matter can neither be created nor destroyed during chemical and physical changes |

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|  | **amorphous solid**        the particles are not arranged in a pattern |

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|  | **boiling**   (boilihNG) *noun,adjective,adverb*    the process that occurs when vaporization takes place inside a liquid as well as on the surface |

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|  | **boiling point**        the temperature at which a liquid vaporizes and changes to a gas |

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|  | **condensation**   (konduhnsayshuhn) *noun*    the changing of a substance from a gas to a liquid |

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|  | **crystalline solid**        the particles form a regular, repeating pattern |

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|  | **evaporation**   (ihvapurayshuhn) *noun,verb*    the process that occurs when vaporization only takes place on the surface of a liquid. |

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|  | **fluid**   (floouhd) *noun,adjective*    a substance that can easily flow |

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|  | **freezing**   (freezihNG) *noun,adjective*    the change from the liquid to the solid state of matter. |

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|  | **freezing point**        temperature at which a liquid changes to a solid |

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|  | **gas**    *noun,verb*    a substance that has neither a definite shape nor a definite volume |

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|  | **liquid**   (lihkwuhd) *noun,adjective*    a substance that has definite volume but no definite shape |

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|  | **melting**   (mEHltihNG) *adjective,noun,verb*    the change from the solid to the liquid state |

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|  | **melting point**        temperature at which a solid changes to a liquid |

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|  | **plasma**   (plazmuh) *noun*    a state of matter that forms when temperatures are high enough to completely remove electrons from the atoms to which they were bound |

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|  | **pressure**   (prEHshur) *noun,verb*    the force of the outward push on a surface divided by the total area over which the force is exerted |

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|  | **solid**   (soluhd) *noun,adjective*    a substance with a definite shape and volume |

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|  | **state of matter**        the physical form in which matter exists; also called phase of matter |

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|  | **sublimation**    *noun,verb*    a phase change from solid directly to gas or from gas directly to solid, w/o ever turning into a liquid |

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|  | **surface tension**        the result of an inward pull among the molecules of a liquid that brings the molecules on the surface closer together. |

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|  | **vaporization**   (vaypuruhzayshuhn) *noun,verb*    process in which a liquid changes to a gas |

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|  | **viscosity**   (vihskosuhtee) *noun*    a liquid’s resistance to flowing  **Unit 2 Atoms Vocabulary** |

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| **atomic mass**        the average mass of all the isotopes of one element |

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|  | **atomic mass unit**        unit of measure that is equal to the mass of one proton |

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|  | **atomic number**        the number of protons in the nucleus of an atom |

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|  | **atoms**   (atuhmz) *noun*    the basic building blocks of all matter |

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|  | **chemical symbol**        the one or two letter name of an element in the periodic table |

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|  | **electrons**   (ihlEHktronz) *noun*    subatomic particles that move and spin around the nucleus and carry a negative (-) charge |

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|  | **element**   (EHluhmuhnt) *noun*    a pure substance that cannot be broken down into simpler substances by any ordinary chemical means |

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|  | **energy level**        the specific amount of energy an electron has |

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|  | **isotopes**   (iisuhtohps)     atoms that have the same number of protons but different number of neutrons |

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|  | **mass number**        the sum of protons and neutrons in the nucleus of an atom |

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|  | **neutral**   (nootruhl) *noun,adjective*    having no overall charge |

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|  | **neutrons**   (nootronz) *noun*    subatomic particles in the nucleus of an atom that have no (0) charge |

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|  | **nucleus**   (nookleeuhs) *noun*    center of an atom made of protons and usually, neutrons |

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|  | **protons**   (prohtonz) *noun*    subatomic particles located in the nucleus of an atom and carry a positive (+) |
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**Unit 2 Periodic Table Vocabulary**

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| **alkali metal**        they are the most reactive metals; their atoms have one electron in their outer level |

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|  | **alkaline earth metals**        they are reactive metals but less reactive than alkali metals; their atoms have two electrons in their outer level |

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|  | **conductivity**   (konduhktihvuhtee)     the ability of an object to transfer heat or electricity to another object |

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|  | **corrosion**   (kurohzhuhn) *noun*    the gradual wearing away of a metal element due to a chemical reaction |

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|  | **ductile**   (duhktuhl) *adjective*    a term used to describe a material that can be pulled into a long wire |

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|  | **group**   (groop) *noun*    elements in the same vertical column of the periodic table; also called a family |

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|  | **halogens**        the elements in group 17 of the periodic table; they are very reactive nonmetals |

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|  | **malleable**   (maleeuhbuhl) *adjective*    a term used to describe a material that can be pounded into shapes |

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|  | **metal**   (mEHtuhl) *noun,verb,adjective*    a class of elements characterized by physical properties that include shininess, malleability, ductility, and conductivity |

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|  | **metalloid**        an element that has some characteristics of both metals and non-metals |

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|  | **nobel gas**        an element in group 18 of the periodic table;they have 8 electrons in their stable outer shell |

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|  | **nonmetal**    *noun,adjective*    an element that lacks most of the properties of a metal |

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|  | **period**   (pihreeuhd) *noun*    a horizontal row of elements in the periodic table |

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|  | **periodic table**        a chart of the elements showing the repeating pattern of their properties |

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|  | **reactivity**   (reeaktihvuhtee)     the ease and speed in which an element combines, or reacts with other elements and compounds |

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|  | **semiconductor**   (sEHmeekuhnduhktur)     a substance that conducts electricity better than an insulator but not as well as a conductor |

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|  | **transition metal**        one of the elements in groups 3-12 of the periodic table |